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pH. pH is defined as the decimal logarithm of the reciprocal of the hydrogen ion activity, a_{H^+} , in a solution.
$$pH = -\log_{10} (a_{H^+}) = \log_{10} \left(\frac{1}{a_{H^+}} \right)$$

~~pH~~ - Wikipedia

$pH = -\log [H^+]$ A water source 's pH value is a function of its acidity, or alkalinity. The pH level is a function of the hydrogen atom activity, as the hydrogen activity is a reasonable indicator of the water's acidity or alkalinity. As seen below, the pH scale ranges from 0 to 14 when 7.0 is neutral.

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~~pH Chemistry (Acids & Bases) — Definition, Calculating pH ...~~

A substance that is neither acidic nor basic is neutral. The pH scale measures how acidic or basic a substance is. The pH scale ranges from 0 to 14. A pH of 7 is neutral.

~~pH Scale — Elmhurst University~~

Each one-unit change in the pH scale corresponds to a ten-fold change in hydrogen ion concentration. The pH scale is theoretically open-ended but most pH values are in the range from 0 to 14. It's a lot easier to use a logarithmic scale instead of always having to write down all those zeros!

~~Acids, Bases, & the pH Scale~~

pH is a measurement of electrically charged particles in a substance. It indicates how acidic or alkaline (basic) that substance is. The pH scale ranges from 0 to 14: Acidic water has a pH lower...

~~pH of Drinking Water: Acceptable Levels and More~~

The term "pH" is a measure of acidity; the lower its value, the more acid the food. The equipment used for determining pH was generally pH meters. Canning methods are described on this page. Some liquids, like lemon juice, lime juice and vinegar, are used as acidifiers, to help lower the pH of foods to increase safety.

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~~Master List of Typical pH and Acid Content of Fruits and ...~~

Soil pH is a measure of the acidity or basicity (alkalinity) of a soil. pH is defined as the negative logarithm (base 10) of the activity of hydronium ions (H^+ or, more precisely, H_3O^+ aq) in a solution.

~~Soil pH - Wikipedia~~

The pH is the measurement of how acidic or alkaline a person's urine is. Doctors often test the urine pH, and they may perform other diagnostic tests, when a person has symptoms that may be related...

~~Urine pH: Normal ranges and what they mean~~

A laboratory will test your urine pH and return results. A neutral pH is 7.0. The higher the number, the more basic (alkaline) it is. The lower the number, the more acidic your urine is.

~~Urine pH Level Test: Purpose, Procedure & Side Effects~~

pH is a logarithmic measure of the hydrogen ion concentration of an aqueous solution $pH = -\log[H^+]$ where \log is the base 10 logarithm and $[H^+]$ is the hydrogen ion concentration in moles per liter pH describes how acidic or basic an aqueous solution is, where a pH below 7 is acidic and a pH greater than 7 is basic. pH of 7 is considered neutral (e.g., pure water).

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~~What Is pH and What Does It Measure? — ThoughtCo~~

Definition of pH. (Entry 1 of 2) : a measure of acidity and alkalinity of a solution that is a number on a scale on which a value of 7 represents neutrality and lower numbers indicate increasing acidity and higher numbers increasing alkalinity and on which each unit of change represents a tenfold change in acidity or alkalinity and that is the negative logarithm of the effective hydrogen-ion concentration or hydrogen-ion activity in gram equivalents per liter of the solution also : the ...

~~Ph | Definition of Ph by Merriam Webster~~

More precisely, pH is defined as the negative log of the hydrogen ion concentration. The range of pH extends from zero to 14. A pH value of 7 is neutral, because pure water has a pH value of exactly 7. Values lower than 7 are acidic; values greater than 7 are basic or alkaline.

~~What is pH? | College of Agriculture, Forestry and Life ...~~

In simple terms, pH is a scale from 1 to 14 that measures the acidity or alkalinity of a liquid. In the middle of the scale is pure distilled water, with a neutral pH of 7. Anything with a pH below...

~~What to know about the pH of water — Medical News Today~~

pH, quantitative measure of the acidity or basicity of aqueous or other liquid solutions. The term, widely used in chemistry, biology, and agronomy, translates

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the values of the concentration of the hydrogen ion into numbers between 0 and 14. Learn more about pH.

~~pH | Definition, Uses, & Facts | Britannica~~

pH. the negative logarithm of the hydrogen ion concentration, $[H^+]$, a measure of the degree to which a solution is acidic or alkaline. An acid is a substance that can give up a hydrogen ion (H^+); a base is a substance that can accept H^+ .

~~pH | definition of pH by Medical dictionary~~

pH. The pH of a solution is a measure of the molar concentration of hydrogen ions in the solution and as such is a measure of the acidity or basicity of the solution. The letters pH stand for "power of hydrogen" and the numerical value is defined as the negative base 10 logarithm of the molar concentration of hydrogen ions. The measurement of the pH of a sample can be done by measuring the cell potential of that sample in reference to a standard hydrogen electrode, as in the accepted ...

~~pH as a Measure of Acid and Base Properties~~

pH is a measure of hydrogen ion concentration, a measure of the acidity or alkalinity of a solution. The pH scale usually ranges from 0 to 14. Aqueous solutions at 25°C with a pH less than 7 are acidic, while those with a pH greater than 7 are basic or alkaline.

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~~pH Definition and Equation in Chemistry — ThoughtCo~~

pH is the term used to refer to the degree of activity of an acid or base (alkali) in the water. It is the most important chemical factor to be maintained in swimming pools. pH is measured on a scale from 0 to 14 with 7 being neutral. Pool water pH is best when kept in the range of 7.2 to 7.8.

~~Pool pH: How to Raise and Lower~~

pH. [pĕ ĩäch] A numerical measure of the acidity or alkalinity of a solution, usually measured on a scale of 0 to 14. Neutral solutions (such as pure water) have a pH of 7, acidic solutions have a pH lower than 7, and alkaline solutions have a pH higher than 7.

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